



Practical handbook for beet-sugar chemists; rapid methods of technico-chemical analyses of the products and by-products and of material used in the manufacture of beet sugar

Werner Moeller-Krause

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...addition of sulphuric acid. Example: Platinum dish and juice 34.628 Dish 24.628 Juice 20.000 Platinum dish with ash 24.737 Dish 24.628 Ash 0.109 Minus 10 per cent, allowance for sulphates 0.0109 Salts 0.0981 This multiplied by 20 and divided by 100 gives 0.3962 per cent. salts contained in the juice. 18.

Determination of the Alkalinity.--The alkalinity of thin juices must be determined frequently and is of great importance in the carbonation and sulphitation process. It is estimated with normal or tenth-normal acid by titration. (For preparation, see Chapter XV.) Phenolphthalein is generally employed as an indicator. In neutral or acid solutions this indicator is colorless but the faintest excess of an alkali produces a bright-red color. A normal acid is a solution of acid in water; and a liter contains the equivalent weight of the active element in grams. This also applies to normal alkali solutions. A normal solution of sulphuric acid therefore contains 49 grams (the molecular weight of sulphuric acid is 98) of the pure acid to the liter, hydrogen being the active element. In this case, the molecular weight must be divided by 2 in order to obtain its equivalent weight: 1. A normal solution of hydrochloric acid contains the full molecular weight (36.5 grams) to the liter, as here only one atom of hydrogen is present. In preparing a normal sodium hydrate solution (NaOH) a liter must contain 40 grams of the salt, the sodium being the active element. The alkalinity is generally determined in grams of CaO per liter of juice. As 1 cubic centimeter of a normal acid solution corresponds to 0.028 gram of CaO, the number of cubic centimeters of acid used to neutralize a certain quantity of juice must be multiplied by 0.028 in order to find the quantity of CaO. It ha...

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